UPSC Chemistry Syllabus 2023 Paper 1

The UPSC Chemistry Syllabus for Paper 1 in 2023 covers the following topics:

UPSC Chemistry Syllabus for Paper 1

Торіс	Sub Topic
Atomic Structure and Chemical Bonding	 Atomic structure, quantum theory, and electronic configurations. Chemical bonding, ionic and covalent bonds, molecular orbital theory.
2. Periodic Table and Periodic Properties	 Periodic trends in properties of elements. Classification of elements, periodicity in properties.
3. Structure and Bonding in Organic Compounds	 Concepts of hybridization, resonance, and electronic effects. Organic reaction mechanisms and stereochemistry.
4. States of Matter	 Gaseous state, laws, and kinetic theory. Liquids and solids, phase equilibria, and phase diagrams.

5. Chemical Thermodynamics	 Laws of thermodynamics, enthalpy, entropy, and free energy. Thermochemistry, Hess's law, and calorimetry.
6. Chemical Kinetics	 Rate of reactions, order, and molecularity. Activation energy, catalysis, and enzyme kinetics.
7. Electrochemistry	 Redox reactions, electrochemical cells, and Nernst equation. Conductance, Kohlrausch's law, and corrosion.
8. Chemical Equilibrium	 Equilibrium constants, Le Chatelier's principle. Acid-base equilibria, pH, and buffers.
9. Surface Chemistry	- Adsorption, colloids, and catalysis.
10. Organic Chemistry	Nomenclature and types of organic reactions.Organic compounds and their functional groups.

- Principles of inorganic chemistry.
- Periodic trends, chemical reactions, and coordination compounds.
- Metals in biological systems, metalloproteins, and metal-based drugs.

11. Inorganic Chemistry

12. Bioinorganic Chemistry

UPSC Chemistry Syllabus 2023 Paper 2

The UPSC Chemistry Syllabus for Paper 2 in 2023 covers a wide range of topics that are essential for a comprehensive understanding of chemistry. The UPSC Chemistry Syllabus for Paper 2 in 2023 covers the following topics:

UPSC Chemistry Syllabus for Mains Paper 2

Topic	Sub Topic
Delocalized Covalent Bonding	General methods (both kinetic and non-kinetic) of study of mechanisms or organic reactions: isotopies, method cross-over experiment, intermediate trapping, stereochemistry; the energy of activation; thermodynamic control and kinetic control of reactions
Reactive intermediates	Generation, geometry, stability, and reactions of carbonium ions and carbanions, free radicals, carbenes, benzynes, and nitrenes.
Substitution reactions	SN1, SN2, and SNi, mechanisms; neighboring group participation; electrophilic and nucleophilic reactions of aromatic compounds including heterocyclic compounds—pyrrole, furan, thiophene, and indole.
Elimination reactions	E1, E2, and E1cb mechanisms; orientation in E2 reactions—Saytzeff and Hoffmann; pyrolytic syn elimination—acetate pyrolysis, Chugaev, and Cope eliminations.

Electrophilic addition to C=C and CC; Addition reactions nucleophilic addition to C=O, CN, conjugated olefins, and carbonyls Pinacol-pinacolone, Hoffmann, Beckmann, Baeyer-Villiger, Favorskii, Fries, Claisen, Reactions and Rearrangement Cope, Stevens, and Wagner—Meerwein rearrangements Classification and examples; Woodward-Hoffmann rules—electrocyclic reactions, Pericyclic reactions cycloaddition reactions [2+2 and 4+2], and sigmatropic shifts [1, 3; 3, 3 and 1, 5], FMO approach Organic polymers polyethylene, polystyrene, Preparation and Properties of Polymers polyvinyl chloride, Teflon, nylon, terylene, synthetic and natural rubber OsO4, HlO4, CrO3, Pb(OAc)4, SeO2, NBS, Synthetic Uses of Reagent B2H6, Na-Liquid NH3, LiAIH4, NaBH4, n-BuLi, MCPBA Photochemical reactions of simple organic compounds, excited and ground states, singlet Photochemistry and triplet states, Norrish-Type I and Type II reactions

Principles and applications in structure elucidation

- (i) Rotational—Diatomic molecules; isotopic substitution and rotational constants.
- (ii) Vibrational—Diatomic molecules, linear triatomic molecules, specific frequencies of functional groups in polyatomic molecules.
- (iii) Electronic—Singlet and triplet states. n and transitions; application to conjugated double bonds and conjugated carbonyls Woodward-Fieser rules; Charge transfer spectra
- (iv) Nuclear Magnetic Resonance (1H NMR): Basic principle; chemical shift and spin-spin interaction and coupling constants. (v) Mass Spectrometry:—Parent peak, base peak, metastable peak, McLafferty rear.

Spectroscopy